

## **IN THE SPECIFICATION**

*Please insert the following paragraph on page 1 after the title of the invention and before the "TECHNICAL FIELD."*

### **--RELATED APPLICATION**

This application is a national phase of PCT/JP2006/300343 filed on January 13, 2006, which claims priority from Japanese Application No. JP 2005-007401 filed on January 14, 2005 and Japanese Application No. JP 2005-377954 filed on December 28, 2005 the disclosures of which Applications are incorporated by reference herein. The benefit of the filing and priority dates of the International and Japanese Applications is respectfully requested.--

*The following paragraphs will replace all prior versions of themselves in the specification of the application.*

- 1) Page 5, lines 4 to 19, please make the following deletions:

### **REFERENCE MARKS IN THE DRAWINGS**

~~1—negative electrode~~

~~1A—collector~~

~~1B—mixture layer~~

~~1C—lead wire~~

~~2—positive electrode~~

~~2A—collector~~

~~2B—mixture layer~~

~~2C—lead wire~~

~~3—separator~~

~~4—battery case~~

~~5 sealing plate~~

~~5A gasket~~

~~5B current break sheet~~

~~6 insulating plate~~

~~10A, 10B, 10C, and 11 point of flection~~

2) Page 8, line 4, please make the following change:

On the other hand, the supporting salt can be of any type of lithium salt without limitations so long as it is dissolvable in the nonaqueous solvent and shows an ionic conductivity. For example, LiPF<sub>6</sub>, LiClO<sub>4</sub>, LiAsF<sub>6</sub>, LiBF<sub>4</sub>, LiB(C<sub>6</sub>H<sub>5</sub>)<sub>4</sub>, LiCH<sub>3</sub>SO<sub>3</sub>, CF<sub>3</sub>SO<sub>3</sub>Li, LiCl, LiN(C<sub>n</sub>C<sub>2n+1</sub>SO<sub>2</sub>)<sub>2</sub> LiN(C<sub>n</sub>C<sub>2n+1</sub>SO<sub>2</sub>)<sub>2</sub>, LiBr, etc. can be used. It is especially desirable to use LiPF<sub>6</sub> as the supporting salt. Any one of these supporting salts may be used singly, or two or more of them may be used as a mixture.

3) Page 11, lines 19 to 20, please make the following change:

Furthermore, it is desirable that one of composites "A" is any of materials expressed as Li<sub>x</sub>Ni<sub>y</sub>Mn<sub>z</sub>Co<sub>1-y-z</sub>O<sub>2</sub> (where 0.9 ≤ x ≤ 0.98, 0.3 ≤ y ≤ 0.4, and 0.3 ≤ z ≤ 0.4) and Li<sub>x</sub>Ni<sub>y</sub>Al<sub>z</sub>Ge<sub>1-y-z</sub>O<sub>2</sub> (where 0.5 ≤ x ≤ 0.85, 0.15 ≤ y ≤ 0.3, and 0.04 ≤ z ≤ 0.17) Li<sub>x</sub>Ni<sub>y</sub>Co<sub>z</sub>Al<sub>1-y-z</sub>O<sub>2</sub> (where 0.9 ≤ x ≤ 0.98, 0.55 ≤ y ≤ 0.8, and 0.15 ≤ z ≤ 0.3). Or, these materials may be combined together. Use of the above materials for composites "A" can further moderate the decreasing rate of discharge voltage near the end of electric discharge, and reduce the heat attributable to the drastic increase of discharge current.

4) Page 17, line 1, please make the following change:

First, evaluations were made for ~~low~~ constant power discharge characteristics with a high load. In an ambient temperature of 20°C, the batteries were charged up to 4.2V with a constant current of 2.6A, and they were further charged with the constant voltage up to the end current of 0.26A. After an interval of 20 minutes, the batteries were discharged by drawing a constant current of 0.52A until their voltages reached 2.0V. This was noted as a first cycle. The batteries were consecutively charged under the same conditions as the first cycle, and after an interval of 20 minutes they were again discharged for a constant output power of 100W until their voltages reached 2.0V. This was noted as a second cycle. During these processes, the batteries were examined for their ratios of discharge capacities in the second cycle with respect to those of the first cycle, as well as their highest temperatures recorded after the end of the second cycle. Examinations were also made at the same time for flection voltages in their discharge voltage curves during discharges of the first cycle. Fig. 2 shows changes in the discharge voltages and the temperatures in the course of electrical discharge of the embodied example 1 and comparison example 1, just for examples.

5) Page 18, line 1, please make the following change:

As shown in Table 1, the batteries represented as the embodied examples 1 and 2 exhibited longer life and less heats at the end of discharges as compared to those of the comparison example 1, as they used composite “A” of  $\text{Li}_{0.95}\text{CoO}_2$  having a high average discharge voltage, added with either one of the composite “B”,  $\text{Li}_x\text{Ni}_y\text{Mn}_z\text{Co}_{1-y-z}\text{O}_2$  and  $\text{Li}_x\text{Ni}_y\text{Al}_z\text{Ce}_{1-y-z}\text{O}_2$

Li<sub>x</sub>Ni<sub>y</sub>Co<sub>z</sub>Al<sub>1-y-z</sub>O<sub>2</sub> having a low average discharge voltage. They also exhibited larger discharge capacities as compared with comparison example 2.

- 6) Page 19, lines 3 to 4 and 13 to 14, please make the following changes:

In the embodied examples 3 through 13, any of the materials expressed as Li<sub>x</sub>CoO<sub>2</sub> (where  $0.9 \leq x \leq 0.98$ ), Li<sub>x</sub>Ni<sub>y</sub>Mn<sub>z</sub>Co<sub>1-y-z</sub>O<sub>2</sub> (where  $0.9 \leq x \leq 0.98$ ,  $0.3 \leq y \leq 0.4$ , and  $0.3 \leq z \leq 0.4$ ), and Li<sub>x</sub>Ni<sub>y</sub>Al<sub>z</sub>Co<sub>1-y-z</sub>O<sub>2</sub> (where  $0.5 \leq x \leq 0.85$ ,  $0.15 \leq y \leq 0.3$ , and  $0.04 \leq z \leq 0.17$ ) Li<sub>x</sub>Ni<sub>y</sub>Co<sub>z</sub>Al<sub>1-y-z</sub>O<sub>2</sub> (where  $0.9 \leq x \leq 0.98$ ,  $0.55 \leq y \leq 0.8$ , and  $0.15 \leq z \leq 0.3$ ) were used as respective composite “A”. These composites “A” were added with LiMnO<sub>2</sub> having the average discharge voltage as low as 2.5 to 3.0V of such an amount that it becomes 2 to 20% of the total amount of the entire active materials by weight. These batteries exhibited longer life and much less heats inside the batteries at the end of discharges as compared to those of the comparison example 1. They also exhibited larger discharge capacities as compared with the comparison example 2. Among those shown above, the embodied examples 10 and 11, which used Li<sub>x</sub>Ni<sub>y</sub>Mn<sub>z</sub>Co<sub>1-y-z</sub>O<sub>2</sub> (where  $0.9 \leq x \leq 0.98$ ,  $0.3 \leq y \leq 0.4$ , and  $0.3 \leq z \leq 0.4$ ) and embodied examples 12 and 13, which used Li<sub>x</sub>Ni<sub>y</sub>Al<sub>z</sub>Co<sub>1-y-z</sub>O<sub>2</sub> (where  $0.5 \leq x \leq 0.85$ ,  $0.15 \leq y \leq 0.3$ , and  $0.04 \leq z \leq 0.17$ ) Li<sub>x</sub>Ni<sub>y</sub>Co<sub>z</sub>Al<sub>1-y-z</sub>O<sub>2</sub> (where  $0.9 \leq x \leq 0.98$ ,  $0.55 \leq y \leq 0.8$ , and  $0.15 \leq z \leq 0.3$ ) as their respective composite “A” exhibited superior performance in both the discharge characteristic with high load output and the charge / discharge cycle characteristic.

- 7) Page 19, line 27 to page 20, line 2, please make the following change:

According to the nonaqueous electrolyte secondary battery in the exemplary embodiment of this invention, the decrease in voltage at the end of electrical discharge is alleviated, and the drastic temperature rise inside the battery due to an increase of the discharge current is alleviated. It is so considered, however, that the comparison example 2 is not suitable for such applications that require a large discharge current because its discharge voltage is too low. The embodied example 4 contained too low an amount of the element Li, which might have resulted in a small capacity due to an excessive irreversible capacity. It is thought that this has inevitably caused an increase in the discharge rate and a slight decrease in the discharge characteristic with the constant output of high load. In the case of comparison example 3, ~~it reached the discharge-end voltage due to a rise in the negative electrode potential within the battery capacity~~ the battery capacity is determined by the battery voltage reaching the discharge-end voltage due to a rise in the negative electrode potential, and it did not provide the effect of voltage control by means of the potential change of positive electrode 2.